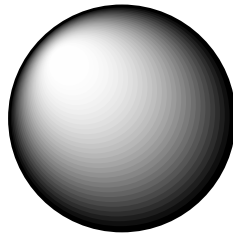


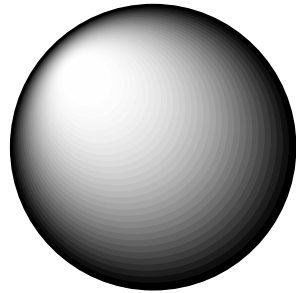
Première période du Tableau Périodique



Orbitale 1s

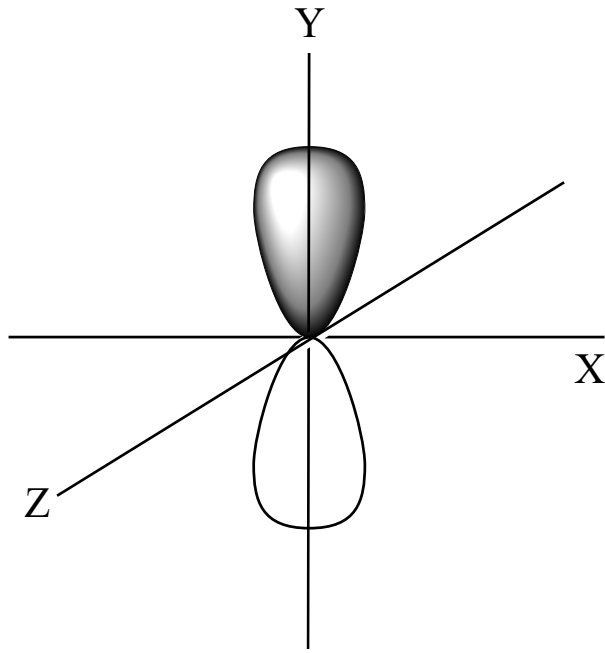
$$n = 1, l = 0, m = 0$$

Deuxième période du Tableau Périodique

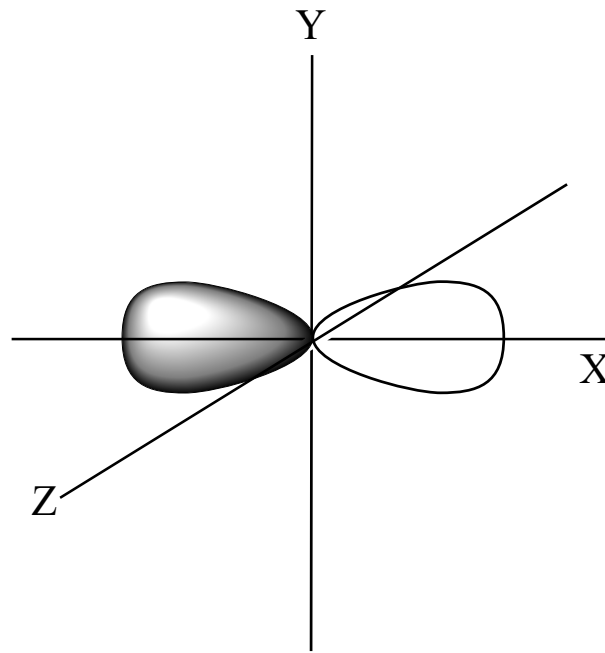


Orbitale 2s

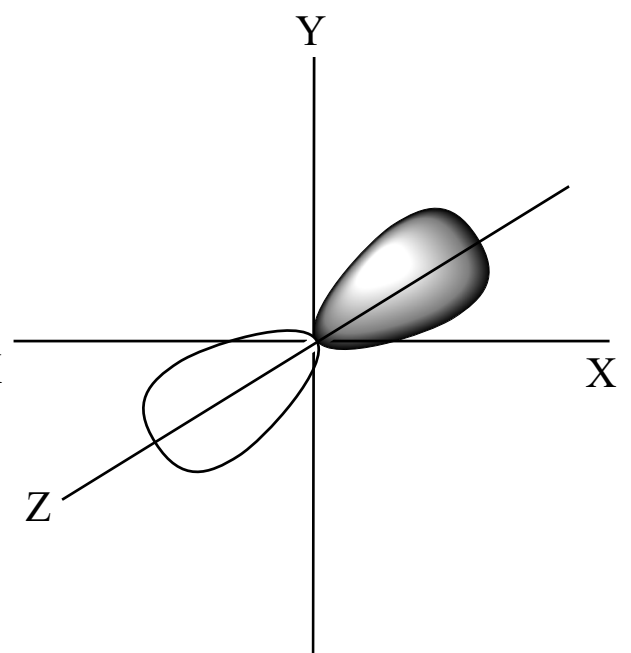
$$n = 2, l = 0, m = 0$$



orbitale 2 p_y



orbitale 2 p_x

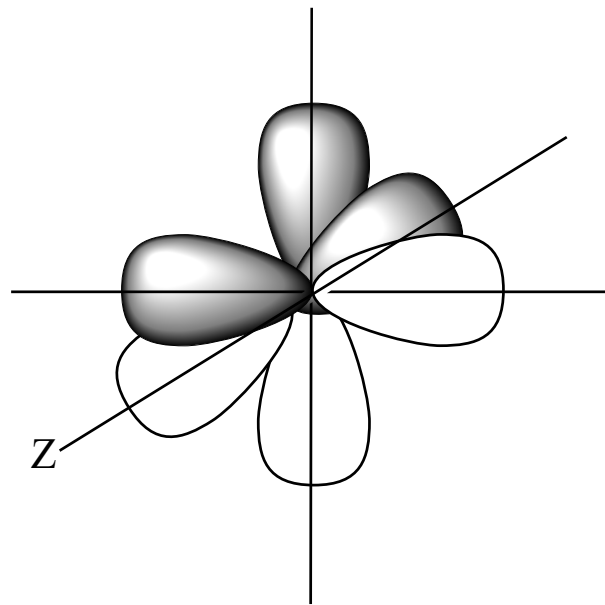


orbitale 2 p_z

$$n = 2, l = 1, m = -1, 0, +1$$

orbitale 2 p_y

Y



X

orbitale 2 p_x

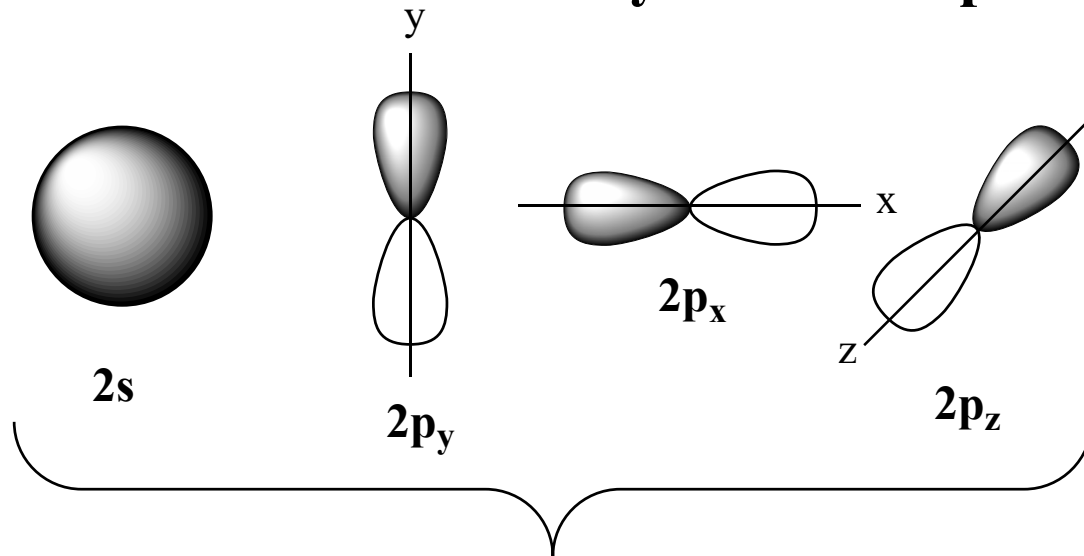
orbitale 2 p_z

Z

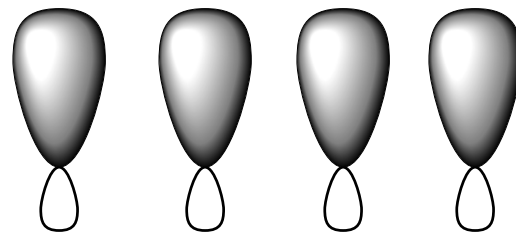
Configuration électronique

1. Les orbitales de basse énergie sont remplies avant celles de haute énergie.
2. Principe d'exclusion de Pauli: une orbitale ne peut contenir plus de deux électrons.
3. Règle de Hund: avant d'être appariés, les électrons sont répartis dans des orbitales vides.

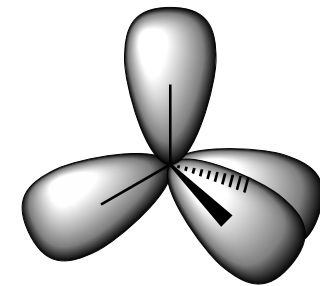
Hybridation sp^3



Hybridation
(combinaison d'orbitales)

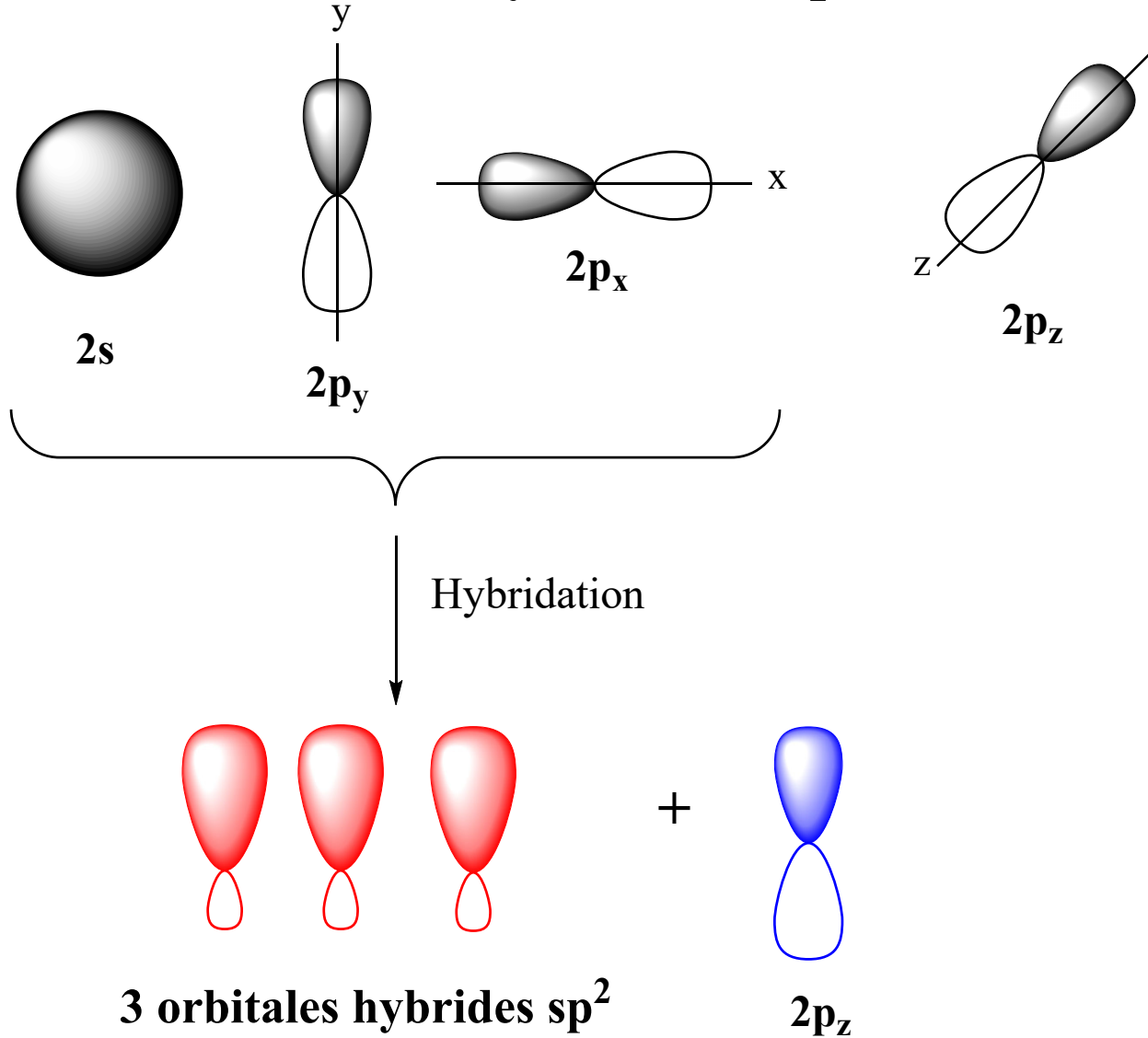


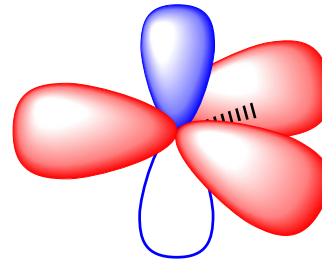
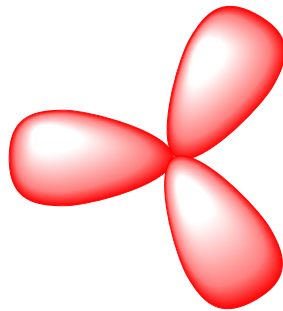
4 orbitales hybrides sp^3



Géométrie tétraédrique

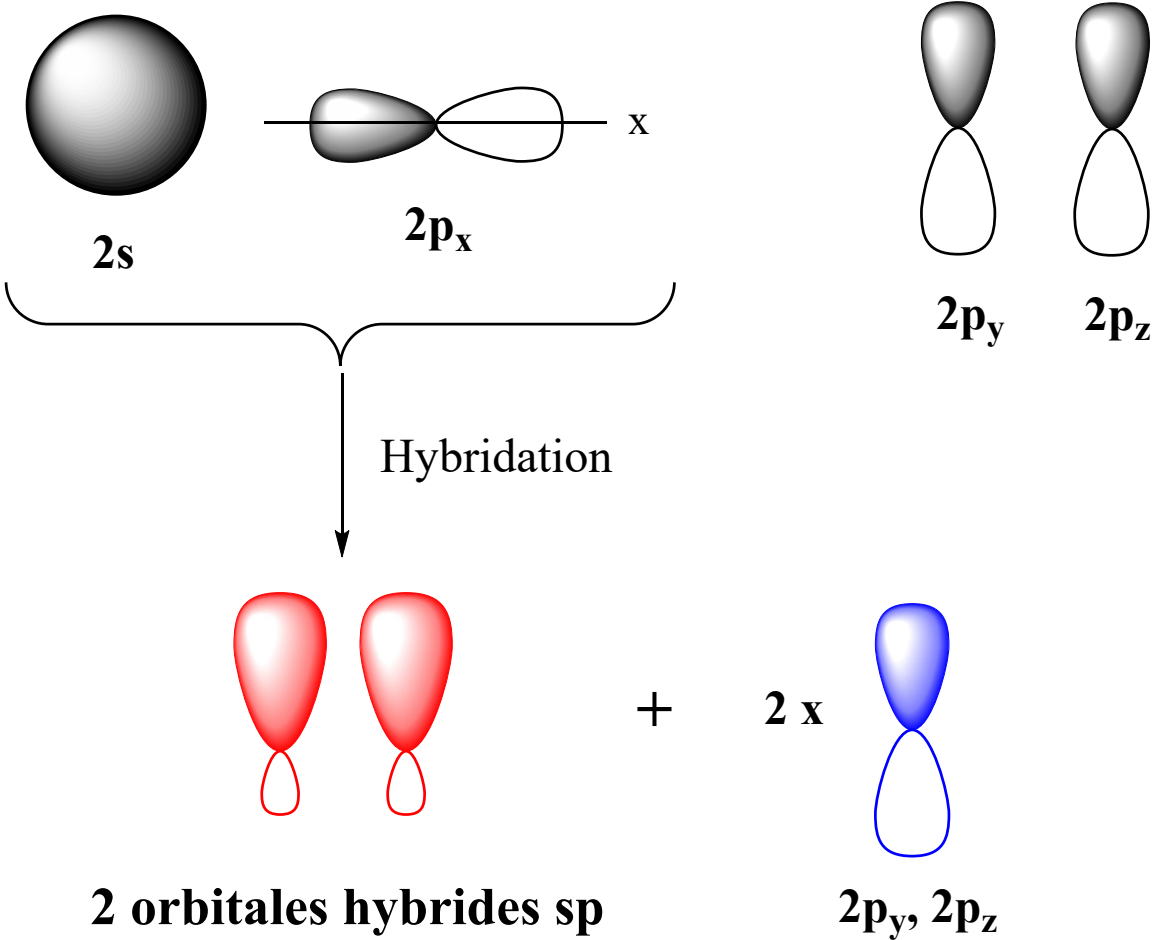
Hybridation sp^2

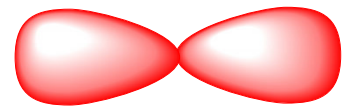




Géométrie trigonale planaire

Hybridation sp





Géométrie linéaire

